Writing ionic equations

# Write balanced ionic equations for the reactions that take place when the following solutions are mixed. Include state symbols in your equations.

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| **1** | Copper chloride solution is mixed with potassium hydroxide solution. |
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| **2** | 2 mL of calcium nitrate solution is added to 2 mL of zinc sulfate solution. |
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| **3** | A few drops of sodium carbonate solution is added to 5 mL of magnesium chloride solution. |
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| **4** | A little potassium iodide solution is poured into a beaker containing lead nitrate solution. |
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| **5** | Ammonia solution is added to a precipitate of zinc hydroxide. |
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| **6** | Limewater (calcium hydroxide solution) is added to aluminium chloride solution. |
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| **7** | Excess sodium hydroxide solution is added to a precipitate of lead hydroxide. |
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| **8** | Iron(II) sulfate solution is added to potassium carbonate solution. |
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| **9** | A few drops of copper chloride solution are added to silver nitrate solution. |
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| **10** | 2 mL of ammonium hydroxide solution is added to 2 mL of iron(III) chloride solution. |
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| **11** | Ammonia solution is added to a precipitate of copper hydroxide. |
|  |  |
| **12** | A few drops of lead nitrate solution are added to iron(II) sulfate solution. |