Writing ionic equations

# Write balanced ionic equations for the reactions that take place when the following solutions are mixed. Include state symbols in your equations.

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| --- | --- | --- | --- | --- | --- | --- |
| **1** | | | Copper chloride solution is mixed with potassium hydroxide solution. | | | |
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| **2** | | | 2 mL of calcium nitrate solution is added to 2 mL of zinc sulfate solution. | | | |
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| **3** | | | A few drops of sodium carbonate solution is added to 5 mL of magnesium chloride solution. | | | |
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| **4** | | | A little potassium iodide solution is poured into a beaker containing lead nitrate solution. | | | |
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| **5** | | | Ammonia solution is added to a precipitate of zinc hydroxide. | |
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| **6** | | | Limewater (calcium hydroxide solution) is added to aluminium chloride solution. | |
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| **7** | | | Excess sodium hydroxide solution is added to a precipitate of lead hydroxide. | |
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| **8** | | | Iron(II) sulfate solution is added to potassium carbonate solution. | |
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| **9** | | | A few drops of copper chloride solution are added to silver nitrate solution. | |
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| **10** | | | 2 mL of ammonium hydroxide solution is added to 2 mL of iron(III) chloride solution. | |
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| **11** | | | Ammonia solution is added to a precipitate of copper hydroxide. | | | |
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| **12** | | | A few drops of lead nitrate solution are added to iron(II) sulfate solution. | | | |