

Dilution calculations

What volumes of the following concentrated solutions are needed to give the stated volumes of the more dilute solutions?

1. 2.00 L of 0.500 molL⁻¹ H₂SO₄ from 2.00 molL⁻¹ H₂SO₄

500 mL

2. 1.00 L of 0.750 molL⁻¹ HCl from 10.00 molL⁻¹ HCl

75 mL

3. 250 mL of 0.250 molL⁻¹ NaOH from 5.00 molL⁻¹ NaOH

12.5 mL

4. 500 mL of 1.25 molL⁻¹ HNO₃ from 5.00 molL⁻¹ HNO₃

125 mL

5. 250 mL of 2.00 molL⁻¹ KOH from 2.50 molL⁻¹ KOH

200 mL