

4B 1b Soluble or insoluble?

Decide whether each of these compounds is soluble, slightly soluble or insoluble.

1	$\text{Fe}(\text{OH})_3$	10	MgCO_3
2	PbSO_4	11	KNO_3
3	AgCl	12	$\text{Cu}(\text{OH})_2$
4	$\text{Pb}(\text{OH})_2$	13	ZnCl_2
5	CaSO_4	14	FeCO_3
6	$(\text{NH}_4)_2\text{CO}_3$	15	KOH
7	$\text{Al}_2(\text{SO}_4)_3$	16	PbCl_2
8	Na_2CO_3	17	$\text{Al}(\text{OH})_3$
9	$\text{Zn}(\text{OH})_2$	18	AgNO_3

4B 1c Predicting precipitates

Will a precipitate form when each of these pairs of solutions are mixed? If so, name it.

1	Iron(II) sulfate solution and lead nitrate solution.
2	Potassium carbonate solution and ammonium hydroxide solution.
3	Zinc chloride solution and aluminium sulfate solution.
4	Magnesium nitrate solution and ammonium carbonate solution.
5	Copper chloride solution and silver nitrate solution.
6	Sodium hydroxide solution and zinc sulfate solution.

4B 1b Answers to: Soluble or insoluble?

1	insoluble	10	insoluble
2	insoluble	11	soluble
3	insoluble	12	insoluble
4	insoluble	13	soluble
5	slightly soluble	14	insoluble
6	soluble	15	soluble
7	soluble	16	slightly soluble
8	soluble	17	insoluble
9	insoluble	18	soluble

4B 1c Answers to: Predicting precipitates

1	Yes: lead sulfate
2	No
3	No
4	Yes: magnesium carbonate
5	Yes: silver chloride
6	Yes: zinc hydroxide